

Tetra- μ_2 -acetato-tetraaquad- μ_3 -oxido-octaoxidotetrauranium(VI) methanol disolvate tetrahydrate

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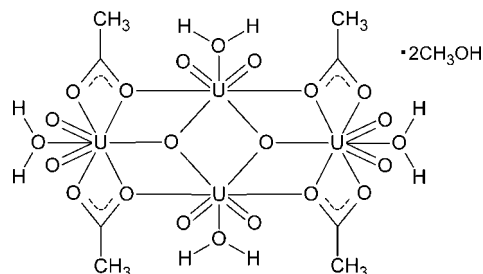
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Key indicators: single-crystal X-ray study; $T = 163$ K; mean $\sigma(\text{C}-\text{C}) = 0.013$ Å; R factor = 0.034; wR factor = 0.088; data-to-parameter ratio = 17.3.

The centrosymmetric title tetramer, $[\text{U}_4(\text{C}_2\text{H}_3\text{O}_2)_4\text{O}_{10}(\text{H}_2\text{O})_4] \cdot 2\text{CH}_4\text{O}$, has a near planar core [maximum deviation from the least squares plane of 0.294 (6) Å]. It consists of two hexagonal-bipyramidally coordinated U^{VI} atoms connected via μ_2 -O (acetate) and μ_3 -O (oxide) bridges in the equatorial plane to two pentagonal-bipyramidally coordinated U^{VI} atoms. The equatorial plane of each U^{VI} atom is completed by a bound water molecule, while the axial positions are occupied by uranyl $(\text{UO}_2)^{2+}$ O atoms. Multiple $\text{O}-\text{H} \cdots \text{O}$ hydrogen bonds are present, including a lattice methanol molecule bound to one of the pentagonal bipyramidal uranyl O atoms, as well as two different $\text{C}_1^1(6)$ chains originating from a donor water molecule, via a uranyl oxygen acceptor and an acetate acceptor on different, adjacent tetramers. Finally, the unit cell contains four U^{VI} tetramers, all connected by hydrogen bonding, forming a supramolecular $R_4^4(24)$ ring.

Related literature

For structurally similar tetrameric complexes with U^{VI} , $M_4[(\text{UO}_2)_4(\mu_3\text{-O})_2L_4]$ ($M = \text{NH}_4^+$, K^+ , Cs^+ ; $L = \text{phthalate}$), see: Charushnikova *et al.* (2005), and with Bi, $[\text{Bi}_2(\mu_3\text{-O})(\text{OCH}(\text{CF}_3)_2)_2(\mu\text{-OCH}(\text{CF}_3)_2)_2(\text{Solv})_2]$ ($\text{Solv} = \text{C}_7\text{H}_8$, Et_2O , thf), see: Andrews *et al.* (2008). For a planar, mixed valent $\text{U}_2^{\text{V}}\text{U}_2^{\text{VI}}$ alkoxide, see: Zozulin *et al.* (1982). For a *p*-benzylcalix[7]arene complex containing a hexanuclear U^{VI} cluster with a planar tetrameric core, see: Thuéry *et al.* (1999), and for dinuclear uranyl-containing salen [*N,N'*-ethylenebis(salicylimine)] complexes, see: Amato *et al.* (2007). For bond-valence-sum calculations, see: Wills (2010).



Experimental

Crystal data

$[\text{U}_4(\text{C}_2\text{H}_3\text{O}_2)_4\text{O}_{10}(\text{H}_2\text{O})_4] \cdot 2\text{CH}_4\text{O}$
 $M_r = 1484.44$
Monoclinic, $P2_1/c$
 $a = 8.334$ (3) Å
 $b = 10.649$ (3) Å
 $c = 16.763$ (5) Å
 $\beta = 107.632$ (4)°

$V = 1417.8$ (8) Å³
 $Z = 2$
Mo $K\alpha$ radiation
 $\mu = 22.87$ mm⁻¹
 $T = 163$ K
0.10 × 0.07 × 0.05 mm

Data collection

Rigaku Saturn70 CCD diffractometer
Absorption correction: numerical (*ABSCOR*; Higashi, 1999)
 $T_{\text{min}} = 0.638$, $T_{\text{max}} = 0.880$

15149 measured reflections
3255 independent reflections
3136 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.076$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.034$
 $wR(F^2) = 0.088$
 $S = 1.09$
3255 reflections
188 parameters
6 restraints

H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\text{max}} = 1.71$ e Å⁻³
 $\Delta\rho_{\text{min}} = -2.85$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-H \cdots A$	$D-H$	$H \cdots A$	$D \cdots A$	$D-H \cdots A$
$\text{O12}-\text{H12} \cdots \text{O9}$	0.84	2.31	3.009 (10)	141
$\text{O8}-\text{H8A} \cdots \text{O7}^i$	0.87 (6)	2.07 (6)	2.859 (8)	151 (7)
$\text{O8}-\text{H8B} \cdots \text{O12}^{ii}$	0.88 (7)	1.78 (6)	2.645 (8)	170 (9)
$\text{O11}-\text{H11A} \cdots \text{O10}^{iii}$	0.87 (7)	1.88 (7)	2.736 (7)	166 (8)
$\text{O11}-\text{H11B} \cdots \text{O2}^{iv}$	0.88 (7)	1.83 (7)	2.705 (7)	173 (7)

Symmetry codes: (i) $-x + 1, y + \frac{1}{2}, -z + \frac{1}{2}$; (ii) $x, -y + \frac{1}{2}, z - \frac{1}{2}$; (iii) $-x + 2, -y, -z + 1$; (iv) $x + 1, -y + \frac{1}{2}, z + \frac{1}{2}$.

Data collection: *CrystalClear-SM Expert* (Rigaku, 2009); cell refinement: *CrystalClear-SM Expert*; data reduction: *CrystalClear-SM Expert*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *Mercury* (Macrae *et al.*, 2006); software used to prepare material for publication: *publCIF* (Westrip, 2010).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BR2182).

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supplementary materials

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Tetra- μ_2 -acetato-tetraaquad- μ_3 -oxido-octaoxidotetrauranium(VI) methanol disolvate tetrahydrate

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Comment

In connection with our on-going studies on metal-binding properties of a new series of macrocyclic polyamide compounds, we were interested in determining whether such macrocycles, by analogy with analogous salen (N,N'-Ethylenebis(salicylimine)) compounds, would also form stable uranyl complexes. Using similar conditions to those which Amato *et al.* (2007) employed with their salens, the title compound crystallized out in the form reported herein.

Bond valence sum calculations (Wills, 2010) performed on the title compound indicate that both crystallographically independent uranium atoms are in the +6 oxidations state ($U1 = 6.095$, $U2 = 6.074$). The centrosymmetric U(VI) tetramer consists of two hexagonal bipyramids, and two pentagonal bipyramids ($U1$ and $U2$, and symmetry equivalents $(-x + 1, -y, -z + 1)$, respectively; Figure 1), with $(UO_2)^{2+}$ oxygen-atoms occupying the axial positions for both $U1$ and $U2$. For $U1$, the equatorial plane consists of *trans*-bidentate acetate anions, each with one bridging μ_2 -O-atom ($O1$ and $O3$), and a μ_3 - O^{2-} anion ($O5$) *trans* to a water molecule. The equatorial plane of $U2$ is therefore composed of the aforementioned μ_2 -O and μ_3 -O atoms, and their inversion-symmetry generated counter-parts, as well as water molecule ($O11$). Similar to the description given by Andrews *et al.* (2008) for $[Bi_2(\mu_3-O)(OCH(CF_3)_2)_2(\mu-OCH(CF_3)_2)_2(Solv)]_2$ ($Solv = C_7H_8, Et_2O, thf$) tetramers, this complex consists of a near planar, ten atom "raft", with maximum deviation from the least squares plane [$U1, U2, O1-5$ and symmetry equivalents $(-x + 1, -y, -z + 1)$] of $0.294(6)$ Å for $O5$. Examination of longer range interactions reveals numerous hydrogen bonds, including a lattice solvent methanol molecule bound to one of the pentagonal bipyramidal uranyl oxygen atoms ($O12-H12\cdots O9$; Figure 1), which further bridges to a bound water molecule of a second tetramer ($O8-H8B\cdots O12^{iii}$, (iii) $x, -y+1/2, z-1/2$; Figure 2, green dashed lines.) The tetramers interact *via* an additional hydrogen bond, whereby the aforementioned water molecule acts as a donor to one of the hexagonal bipyramidal uranyl oxygen atoms on the first assembly ($O8-H8A\cdots O7^{ii}$, (ii) $-x+1, y+1/2, -z+1/2$; Figure 2, black dashed lines) leading to a 24-membered H-bonded ring (graph set notation $R^4_4(24)$) which spans all four tetramers present in the unit cell.

Experimental

160 mg (0.337 mmol) of uranyl acetate dihydrate ($UO_2(CH_3COO)_2 \cdot 2H_2O$) was dissolved in 1 ml methanol, and warmed in a hot water bath (323 K) for 5 minutes. The solution was left at room temperature (293 K) for slow evaporation. Yellow prismatic crystals, suitable for analysis by X-ray diffraction, formed after one week.

Refinement

The water H-atoms, $H8(A,B)$ and $H11(A,B)$, were located from difference Fourier maps, and were refined with distance and angle restraints: $O-H = 0.87(2)$ Å, $H-O-H = 104.50(4)^\circ$. The C-bound methyl and methanolic H-atoms were included in calculated positions and treated as riding atoms: $X-H = 0.98$ and 0.84 Å for *H*-methyl and *H*-OMe H-atoms, respectively.

supplementary materials

For all H atoms, $U_{\text{iso}}(\text{H}) = k \times U_{\text{eq}}(\text{parent atom})$, where $k = 1.2$ for *H*-methyl and $k = 1.5$ for all O-bound H-atoms. The final electron density synthesis shows the highest peak of $1.71 \text{ e}\text{\AA}^3$ located 0.98 \AA from U2 and the deepest hole of $-2.85 \text{ e}\text{\AA}^3$ located 0.72 \AA from U1 and may be the result of residual absorption effects.

Figures

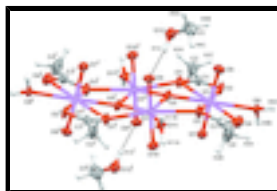


Fig. 1. 50% displacement ellipsoid representation of the title compound; hydrogen bonding indicated by dashed lines. Symmetry code: (i) $-x + 1, -y, -z + 1$.

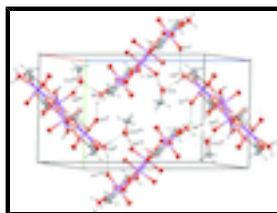


Fig. 2. 50% displacement ellipsoid representation of the packed unit cell for the title compound; hydrogen bonding indicated by dashed lines. H-bonding for $\text{O11}—\text{H11A}\cdots\text{O10}^{\text{iv}}$ and $\text{O11}—\text{H11B}\cdots\text{O2}^{\text{v}}$ not shown. Symmetry codes: (ii) $-x + 1, y + 1/2, -z + 1/2$, (vi) $x, -y + 1/2, z + 1/2$, (vii) $-x + 1, -y + 1, -z + 1$, (viii) $x, y + 1, z$, (ix) $-x + 1, y + 1/2, -z + 3/2$.

Tetra- μ_2 -acetato-tetraaquadi- μ_3 -oxido-octaoxidotetrauranium(VI) methanol disolvate tetrahydrate

Crystal data

$[\text{U}_4(\text{C}_2\text{H}_3\text{O}_2)_4\text{O}_{10}(\text{H}_2\text{O})_4]\cdot 2\text{CH}_4\text{O}$

$M_r = 1484.44$

Monoclinic, $P2_1/c$

Hall symbol: $-P\ 2ybc$

$a = 8.334\ (3)\ \text{\AA}$

$b = 10.649\ (3)\ \text{\AA}$

$c = 16.763\ (5)\ \text{\AA}$

$\beta = 107.632\ (4)^\circ$

$V = 1417.8\ (8)\ \text{\AA}^3$

$Z = 2$

$F(000) = 1296$

$D_x = 3.477\ \text{Mg m}^{-3}$

Mo $K\alpha$ radiation, $\lambda = 0.71075\ \text{\AA}$

Cell parameters from 4946 reflections

$\theta = 1.9\text{--}29.5^\circ$

$\mu = 22.87\ \text{mm}^{-1}$

$T = 163\ \text{K}$

Prism, yellow

$0.10 \times 0.07 \times 0.05\ \text{mm}$

Data collection

Rigaku Saturn70 CCD
diffractometer

Radiation source: fine-focus sealed tube
graphite - Rigaku SHINE

Detector resolution: $14.63\ \text{pixels mm}^{-1}$

ω scans

Absorption correction: numerical
(*ABSCOR*; Higashi, 1999)

$T_{\text{min}} = 0.638, T_{\text{max}} = 0.880$

15149 measured reflections

3255 independent reflections

3136 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.076$

$\theta_{\text{max}} = 27.5^\circ, \theta_{\text{min}} = 3.1^\circ$

$h = -10 \rightarrow 10$

$k = -13 \rightarrow 13$

$l = -21 \rightarrow 21$

Refinement

Refinement on F^2	Primary atom site location: structure-invariant direct methods
Least-squares matrix: full	Secondary atom site location: difference Fourier map
$R[F^2 > 2\sigma(F^2)] = 0.034$	Hydrogen site location: inferred from neighbouring sites
$wR(F^2) = 0.088$	H atoms treated by a mixture of independent and constrained refinement
$S = 1.09$	$w = 1/[\sigma^2(F_o^2) + (0.0469P)^2 + 6.6005P]$
3255 reflections	where $P = (F_o^2 + 2F_c^2)/3$
188 parameters	$(\Delta/\sigma)_{\max} = 0.001$
6 restraints	$\Delta\rho_{\max} = 1.71 \text{ e } \text{\AA}^{-3}$
	$\Delta\rho_{\min} = -2.85 \text{ e } \text{\AA}^{-3}$

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
U1	0.41912 (3)	0.20869 (2)	0.332933 (15)	0.01529 (10)
U2	0.71490 (3)	0.06455 (2)	0.542406 (14)	0.01236 (10)
O1	0.1576 (7)	0.0796 (5)	0.3408 (3)	0.0247 (11)
O2	0.1274 (7)	0.1931 (5)	0.2295 (3)	0.0272 (12)
O3	0.7084 (7)	0.2370 (6)	0.4457 (4)	0.0296 (13)
O4	0.6738 (7)	0.3262 (6)	0.3266 (4)	0.0306 (13)
O5	0.4581 (6)	0.0936 (5)	0.4484 (3)	0.0217 (11)
O6	0.3486 (7)	0.3399 (5)	0.3788 (4)	0.0277 (12)
O7	0.4880 (7)	0.0816 (5)	0.2841 (4)	0.0277 (12)
O8	0.3755 (7)	0.3336 (5)	0.2040 (3)	0.0230 (11)
H8A	0.403 (12)	0.406 (4)	0.188 (5)	0.034*
H8B	0.367 (13)	0.286 (6)	0.160 (4)	0.034*
O9	0.6588 (7)	0.1628 (5)	0.6159 (4)	0.0264 (12)
O10	0.7947 (6)	-0.0283 (5)	0.4734 (3)	0.0236 (11)
O11	0.9960 (6)	0.1499 (5)	0.5992 (3)	0.0202 (10)
H11A	1.075 (7)	0.121 (8)	0.580 (5)	0.030*
H11B	1.047 (9)	0.200 (7)	0.641 (4)	0.030*

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O12	0.3208 (9)	0.2904 (6)	0.5615 (4)	0.0376 (15)
H12	0.3876	0.2346	0.5557	0.056*
C1	0.0629 (9)	0.1183 (7)	0.2696 (4)	0.0207 (14)
C2	-0.1168 (10)	0.0795 (8)	0.2383 (6)	0.0332 (19)
H2A	-0.1601	0.0663	0.2859	0.040*
H2B	-0.1261	0.0013	0.2064	0.040*
H2C	-0.1825	0.1454	0.2020	0.040*
C3	0.7675 (10)	0.2993 (7)	0.3966 (5)	0.0239 (16)
C4	0.9472 (12)	0.3420 (11)	0.4242 (7)	0.049 (3)
H4A	0.9639	0.4053	0.3850	0.059*
H4B	1.0212	0.2701	0.4252	0.059*
H4C	0.9741	0.3785	0.4803	0.059*
C5	0.4006 (16)	0.4061 (10)	0.5715 (6)	0.051 (3)
H5A	0.4034	0.4418	0.6258	0.062*
H5B	0.3388	0.4628	0.5267	0.062*
H5C	0.5159	0.3955	0.5692	0.062*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
U1	0.01397 (15)	0.01724 (15)	0.01456 (16)	0.00003 (8)	0.00417 (11)	0.00355 (8)
U2	0.00972 (14)	0.01612 (15)	0.01148 (15)	-0.00104 (8)	0.00357 (10)	0.00029 (8)
O1	0.018 (3)	0.031 (3)	0.021 (3)	-0.001 (2)	0.000 (2)	0.014 (2)
O2	0.023 (3)	0.037 (3)	0.018 (3)	-0.006 (2)	0.002 (2)	0.015 (2)
O3	0.020 (3)	0.034 (3)	0.033 (3)	-0.006 (2)	0.005 (2)	0.016 (3)
O4	0.022 (3)	0.043 (3)	0.025 (3)	-0.007 (2)	0.003 (2)	0.005 (2)
O5	0.012 (2)	0.031 (3)	0.019 (3)	-0.006 (2)	0.0002 (19)	0.010 (2)
O6	0.019 (3)	0.032 (3)	0.029 (3)	0.006 (2)	0.003 (2)	-0.001 (2)
O7	0.026 (3)	0.029 (3)	0.029 (3)	0.002 (2)	0.011 (2)	-0.001 (2)
O8	0.023 (3)	0.025 (3)	0.022 (3)	0.001 (2)	0.008 (2)	0.011 (2)
O9	0.021 (3)	0.029 (3)	0.032 (3)	0.000 (2)	0.011 (2)	-0.008 (2)
O10	0.016 (2)	0.032 (3)	0.022 (3)	-0.004 (2)	0.004 (2)	-0.009 (2)
O11	0.014 (2)	0.027 (3)	0.022 (3)	-0.006 (2)	0.009 (2)	-0.010 (2)
O12	0.042 (4)	0.040 (4)	0.033 (3)	0.004 (3)	0.014 (3)	0.004 (3)
C1	0.017 (3)	0.020 (3)	0.020 (3)	0.000 (3)	-0.001 (3)	0.006 (3)
C2	0.016 (4)	0.043 (5)	0.036 (5)	-0.006 (3)	0.001 (3)	0.014 (4)
C3	0.023 (4)	0.021 (3)	0.031 (4)	-0.005 (3)	0.012 (3)	0.003 (3)
C4	0.028 (5)	0.073 (7)	0.044 (6)	-0.014 (5)	0.006 (4)	0.021 (5)
C5	0.072 (8)	0.041 (5)	0.032 (5)	0.009 (5)	0.002 (5)	-0.005 (4)

Geometric parameters (\AA , $^\circ$)

U1—O7	1.765 (6)	O4—C3	1.230 (10)
U1—O6	1.779 (6)	O5—U2 ⁱ	2.252 (5)
U1—O5	2.230 (5)	O8—H8A	0.87 (2)
U1—O8	2.469 (5)	O8—H8B	0.87 (2)
U1—O4	2.494 (6)	O11—H11A	0.87 (2)
U1—O2	2.528 (6)	O11—H11B	0.87 (2)

U1—O3	2.592 (6)	O12—C5	1.385 (12)
U1—O1	2.614 (5)	O12—H12	0.8400
U2—O9	1.784 (5)	C1—C2	1.487 (10)
U2—O10	1.795 (5)	C2—H2A	0.9800
U2—O5 ⁱ	2.252 (5)	C2—H2B	0.9800
U2—O5	2.261 (5)	C2—H2C	0.9800
U2—O11	2.422 (5)	C3—C4	1.498 (12)
U2—O3	2.438 (5)	C4—H4A	0.9800
U2—O1 ⁱ	2.461 (5)	C4—H4B	0.9800
O1—C1	1.284 (8)	C4—H4C	0.9800
O1—U2 ⁱ	2.461 (5)	C5—H5A	0.9800
O2—C1	1.262 (9)	C5—H5B	0.9800
O3—C3	1.268 (9)	C5—H5C	0.9800
O7—U1—O6	178.0 (3)	O5—U2—O1 ⁱ	136.78 (17)
O7—U1—O5	89.9 (2)	O11—U2—O1 ⁱ	77.77 (18)
O6—U1—O5	92.1 (2)	O3—U2—O1 ⁱ	156.28 (18)
O7—U1—O8	89.4 (2)	C1—O1—U2 ⁱ	159.4 (5)
O6—U1—O8	88.6 (2)	C1—O1—U1	94.1 (4)
O5—U1—O8	179.2 (2)	U2 ⁱ —O1—U1	101.70 (18)
O7—U1—O4	88.0 (2)	C1—O2—U1	98.8 (4)
O6—U1—O4	91.2 (2)	C3—O3—U2	153.0 (5)
O5—U1—O4	114.24 (18)	C3—O3—U1	92.6 (5)
O8—U1—O4	66.05 (18)	U2—O3—U1	103.06 (19)
O7—U1—O2	90.7 (2)	C3—O4—U1	98.3 (5)
O6—U1—O2	88.5 (2)	U1—O5—U2 ⁱ	122.8 (2)
O5—U1—O2	114.60 (17)	U1—O5—U2	122.6 (2)
O8—U1—O2	65.10 (17)	U2 ⁱ —O5—U2	109.9 (2)
O4—U1—O2	131.14 (18)	U1—O8—H8A	140 (6)
O7—U1—O3	93.8 (2)	U1—O8—H8B	112 (5)
O6—U1—O3	87.1 (2)	H8A—O8—H8B	102 (4)
O5—U1—O3	64.64 (17)	U2—O11—H11A	118 (5)
O8—U1—O3	115.72 (17)	U2—O11—H11B	136 (5)
O4—U1—O3	50.00 (18)	H11A—O11—H11B	105 (4)
O2—U1—O3	175.5 (2)	C5—O12—H12	109.5
O7—U1—O1	90.8 (2)	O2—C1—O1	117.2 (6)
O6—U1—O1	90.1 (2)	O2—C1—C2	122.2 (7)
O5—U1—O1	64.63 (17)	O1—C1—C2	120.5 (7)
O8—U1—O1	115.07 (17)	C1—C2—H2A	109.5
O4—U1—O1	178.31 (18)	C1—C2—H2B	109.5
O2—U1—O1	49.97 (16)	H2A—C2—H2B	109.5
O3—U1—O1	129.03 (16)	C1—C2—H2C	109.5
O9—U2—O10	173.7 (2)	H2A—C2—H2C	109.5
O9—U2—O5 ⁱ	94.8 (2)	H2B—C2—H2C	109.5
O10—U2—O5 ⁱ	90.1 (2)	O4—C3—O3	118.9 (7)
O9—U2—O5	90.6 (2)	O4—C3—C4	120.7 (7)
O10—U2—O5	94.8 (2)	O3—C3—C4	120.4 (8)

supplementary materials

O5 ⁱ —U2—O5	70.1 (2)	C3—C4—H4A	109.5
O9—U2—O11	86.2 (2)	C3—C4—H4B	109.5
O10—U2—O11	87.5 (2)	H4A—C4—H4B	109.5
O5 ⁱ —U2—O11	144.73 (18)	C3—C4—H4C	109.5
O5—U2—O11	145.17 (18)	H4A—C4—H4C	109.5
O9—U2—O3	93.4 (2)	H4B—C4—H4C	109.5
O10—U2—O3	85.7 (2)	O12—C5—H5A	109.5
O5 ⁱ —U2—O3	136.25 (19)	O12—C5—H5B	109.5
O5—U2—O3	66.93 (18)	H5A—C5—H5B	109.5
O11—U2—O3	78.65 (18)	O12—C5—H5C	109.5
O9—U2—O1 ⁱ	87.5 (2)	H5A—C5—H5C	109.5
O10—U2—O1 ⁱ	90.8 (2)	H5B—C5—H5C	109.5
O5 ⁱ —U2—O1 ⁱ	67.08 (17)		

Symmetry codes: (i) $-x+1, -y, -z+1$.

Hydrogen-bond geometry (\AA , $^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O12—H12 \cdots O9	0.84	2.31	3.009 (10)	141
O12—H12 \cdots O5	0.84	2.54	3.257 (9)	143
O8—H8A \cdots O7 ⁱⁱ	0.87 (6)	2.07 (6)	2.859 (8)	151 (7)
O8—H8B \cdots O12 ⁱⁱⁱ	0.88 (7)	1.78 (6)	2.645 (8)	170 (9)
O11—H11A \cdots O10 ^{iv}	0.87 (7)	1.88 (7)	2.736 (7)	166 (8)
O11—H11B \cdots O2 ^v	0.88 (7)	1.83 (7)	2.705 (7)	173 (7)

Symmetry codes: (ii) $-x+1, y+1/2, -z+1/2$; (iii) $x, -y+1/2, z-1/2$; (iv) $-x+2, -y, -z+1$; (v) $x+1, -y+1/2, z+1/2$.

Fig. 1

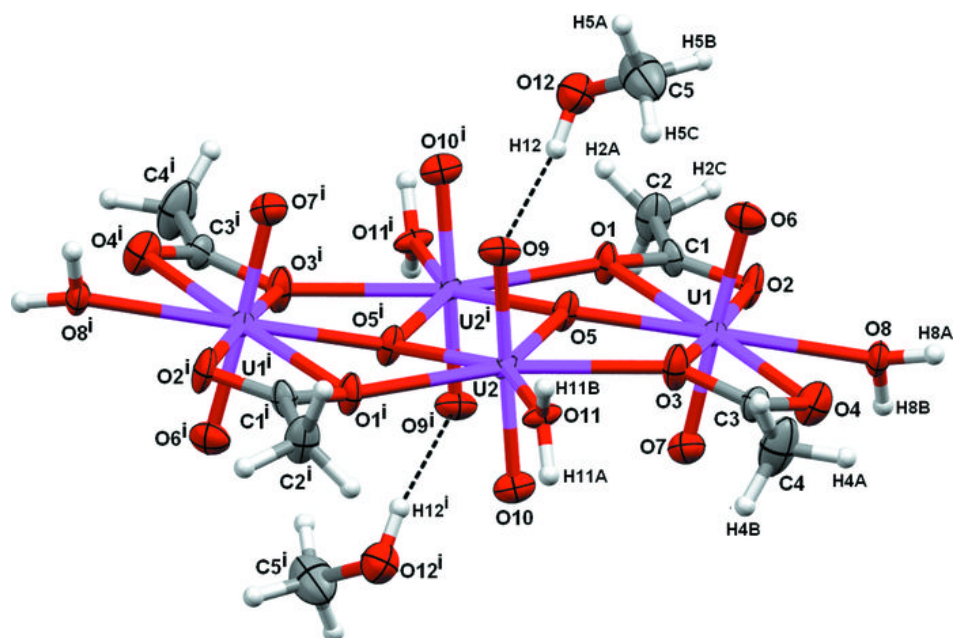


Fig. 2

